UV/TiO2 photodegradation of metronidazole, ciprofloxacin and sulfamethoxazole in aqueous solution: An optimization and kinetic study

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**Table S1** MRM settings, including quantifier, qualifier and CE (collision energy) for all analytes.

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| --- | --- | --- | --- | --- | --- | --- |
| Analyte  | Precursor ion [m/z]  | Quantifier  | Qualifier  | Fragmentor | Dwell voltage | cell acceleratorvoltage |
|   |   | Product ion [m/z]  | CE [V] | Product ion [m/z]  | CE [V] |   | [V] | [V] |
| Metronidazole | 172 | 128 | 18 | 82 | 25 | 230 | 200 | 4 |
| Ciprofloxacin | 332 | 314 | 20 | 231 | 35 | 110 | 200 | 4 |
| Sulfamethoxazole | 254 | 156 | 16 | 92 | 28 | 134 | 200 | 4 |

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| Figure S1 Evolution of the UV–vis absorption spectrum of photocatalytic degradation of SMX (5 mg/L) using TiO2 (0.7g/L) |
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Figure S2 Evolution of the UV–vis absorption spectrum of MNZ (80 mg/L) depending on the irradiation time.



Figure S3 Evolution of the UV–vis absorption spectrum of CIP (80 mg/L) depending on the irradiation time



FigureS4 Evolution of the UV–vis absorption spectrum of (a) [MNZ]=[CIP]=40mg/L, (b)[SMX]=[CIP]=40mg/L, (c)[SMX]=[MNZ]=40mg/L, (d) [MNZ]=40mg/L, [SMX]=[CIP]=20mg/L and (e) [MNZ]= [CIP]= 20mg/L, [SMX]= 40mg/L depending on the irradiation time (TiO2 = 0.7 g/L)